

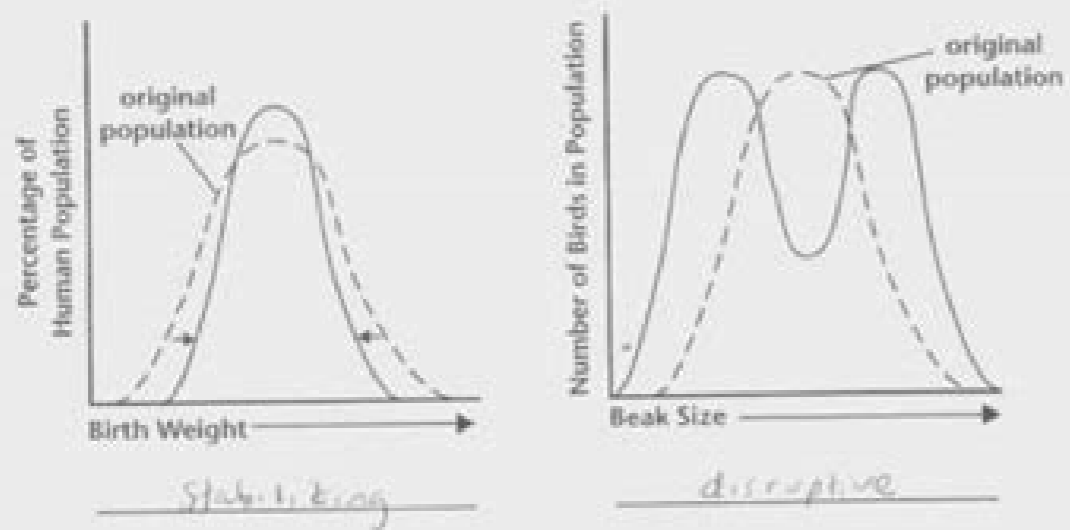
Honors chemistry vsepr worksheet 1 answers

1

Stabilizing and Disruptive Selection

In most populations, a trait that has higher fitness leads to greater numbers of organisms with that trait. On the graphs, dotted lines represent the original population. The solid lines represent the population after selection has taken place.

Identify whether each graph shows stabilizing selection or disruptive selection. Write the type of selection shown below each graph.



Use the graphs to answer the questions.

1. Under which type of selection do organisms in the middle of the curve have the highest fitness? Circle the correct answer.

disruptive stabilizing

2. In disruptive selection, organisms represented by which part of the curve have the lowest fitness? Circle the correct answer.

middle of the curve ends of the curve

3. Describe a situation that might lead to the changes shown in the graph on the right.

environment changes to have 2 different areas

CHM134 GENERAL CHEMISTRY I VSEPR worksheet #1

Complete the following table

Example	# valence electrons	Lewis Structure	Electronic Geometry	Molecular Geometry	a. o. hybrid.	Draw (angles)
CB ₄	32		Td	Td	sp ³	
ClF ₃	28		Tbp	T shaped	sp ³ d	
BCl ₃	24		Trig. Pl.	Trig. Pl.	sp ²	
KrF ₂	22		Tbp	linear	sp ³ d	
SF ₄	34		Tbp	see-saw	sp ³ d	
PCl ₆ ⁻	48		Oh	Oh	sp ³ d ²	

Chemistry Stoichiometry Reflect #2
Learning Target: I can use stoichiometry to calculate amounts

1. Balance the equation
 $2 \text{KNO}_3 \rightarrow 2 \text{KNO}_2 + \text{O}_2$

2. Determine the mass (in grams) of potassium nitrate used if the reaction produced 55.2 g of O₂.
 $55.2 \text{ g O}_2 \left(\frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \right) \left(\frac{2 \text{ mol KNO}_3}{1 \text{ mol O}_2} \right) \left(\frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} \right) = 346.8 \text{ g KNO}_3$

3. Determine the amount of water (in grams) produced in the reaction.
 $100.0 \text{ g H}_2 \left(\frac{1 \text{ mol H}_2}{2.016 \text{ g H}_2} \right) \left(\frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol H}_2} \right) \left(\frac{18.015 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \right) = 1800 \text{ g H}_2\text{O}$

4. Determine how many grams of ammonia (in grams) produced in the reaction.
 $100.0 \text{ g N}_2 \left(\frac{1 \text{ mol N}_2}{28.014 \text{ g N}_2} \right) \left(\frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \right) \left(\frac{17.031 \text{ g NH}_3}{1 \text{ mol NH}_3} \right) = 121.6 \text{ g NH}_3$

5. Using the following chemical reaction and knowing that 1.0 g of copper was produced in this experiment, how many molecules of mercury atoms are also produced?
 $6 \text{Hg} + 2 \text{Cu}_2\text{S} \rightarrow 4 \text{Cu} + 2 \text{Hg}_2\text{S} + \text{S}_2$
 $1.0 \text{ g Cu} \left(\frac{1 \text{ mol Cu}}{63.546 \text{ g Cu}} \right) \left(\frac{2 \text{ mol Hg}_2\text{S}}{4 \text{ mol Cu}} \right) \left(\frac{2 \text{ mol Hg}}{1 \text{ mol Hg}_2\text{S}} \right) \left(\frac{1 \text{ molecule Hg}}{1 \text{ mol Hg}} \right) = 1.57 \times 10^{23} \text{ molecules Hg}$

6. Challenge
 What is the percent yield of the following reaction if 1.0 g of copper (II) chloride was used in the lab and 0.75 g of copper atoms are recovered at the end of the experiment? (Note: you must calculate the theoretical yield of copper atoms first)
 $1.0 \text{ g CuCl}_2 \left(\frac{1 \text{ mol CuCl}_2}{134.45 \text{ g CuCl}_2} \right) \left(\frac{1 \text{ mol Cu}}{1 \text{ mol CuCl}_2} \right) \left(\frac{63.546 \text{ g Cu}}{1 \text{ mol Cu}} \right) = 0.473 \text{ g Cu}$
 $\frac{0.75 \text{ g Cu}}{0.473 \text{ g Cu}} \times 100\% = 158\%$

BLM 2.1.8B

VSEPR Theory Summary Chart

- valence electron pairs – both bonding and lone pairs will orient themselves as far apart as possible around a central atom
- treat multiple bonding pairs (double and triple bonds) as if they were 1 bonding pair for the purposes of counting lone pairs and bonding pairs

VSEPR Class	Name of molecular shape	Types of electron pairs	Example	Lewis dot diagram	VSEPR Diagram	Model
AX ₄	tetrahedral	all BP (0 LP, 4 BP)	CH ₄			
AX ₃ E	tetrahedral based: pyramidal	1 LP, 3 BP	NH ₃			
AX ₂ E ₂	tetrahedral based: v-shaped	2 LP, 2 BP	H ₂ O			
AX ₃	trigonal planar	all BP (0 LP, 3 BP)	CH ₂ O			
AX ₂ E	trigonal planar based: v-shaped	1 LP, 2 BP	SO ₂			
AX ₂	linear	all BP (0 LP, 2 BP)	CO ₂			

X = bonding pair
E = lone pair

NAME _____

INSTRUCTIONS: Write E in the blank if the material is heterogeneous or H if it is homogeneous.

- | | |
|--------------------------------------|-------------------------------------|
| 1. Wood _____ | 6. Dirt _____ |
| 2. Freshly-brewed black coffee _____ | 7. Sausage-and-mushroom pizza _____ |
| 3. Water _____ | 8. Air _____ |
| 4. Lucky Charms® _____ | 9. Milk _____ |
| 5. Salt _____ | 10. Gold _____ |

INSTRUCTIONS: Classify each of the following as an element [E], a compound [C], or a mixture [M].

- | | |
|------------------------------|--------------------------|
| 11. Gold _____ | 16. Air _____ |
| 12. Water _____ | 17. Carbon dioxide _____ |
| 13. Seawater _____ | 18. Silver _____ |
| 14. Sugar _____ | 19. Ice _____ |
| 15. A chocolate sundae _____ | 20. A Big Mac® _____ |

INSTRUCTIONS: Classify each of the following properties of matter as physical [P] or chemical [C].

- | | |
|------------------------------------|--|
| 21. Color _____ | 26. Reacts violently with chlorine _____ |
| 22. Density _____ | 27. Good conductor of heat _____ |
| 23. Burns easily (flammable) _____ | 28. Dissolves readily in water _____ |
| 24. Not affected by acids _____ | 29. Melts at 145 °C _____ |
| 25. Boils at 450 °C _____ | 30. Malleable _____ |

INSTRUCTIONS: Classify each of the following changes in matter as physical [P] or chemical [C].

- | | |
|---------------------------------------|--------------------------------------|
| 31. Grinding chalk into powder _____ | 36. Burning gasoline _____ |
| 32. Dissolving salt in water _____ | 37. Hammering gold into foil _____ |
| 33. Dissolving zinc in acid _____ | 38. Melting ice _____ |
| 34. Tearing a piece of paper _____ | 39. Digesting food _____ |
| 35. Stretching copper into wire _____ | 40. Making hydrogen from water _____ |

INSTRUCTIONS: Classify each of the following as an intensive property [I] or an extensive property [E].

- | | |
|-------------------------|------------------|
| 41. Mass _____ | 46. Color _____ |
| 42. Density _____ | 47. Volume _____ |
| 43. Melting point _____ | 48. Length _____ |

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